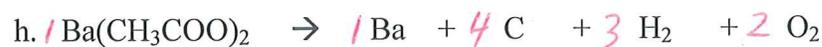
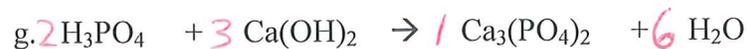
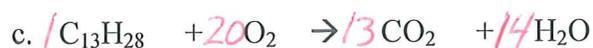


Chemistry 11

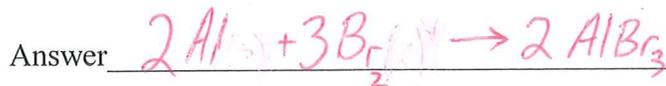
Chemical Equation Balancing

1. Balance the following equations.

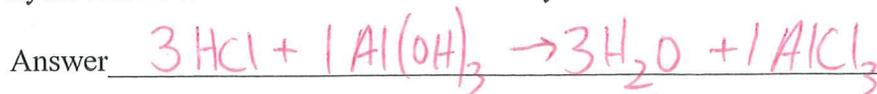


2. Write a balanced chemical equation for each of the following. Don't forget *diatomic* elements! (2 marks each = 24 marks)

a. aluminum metal reacts with bromine to form aluminum bromide

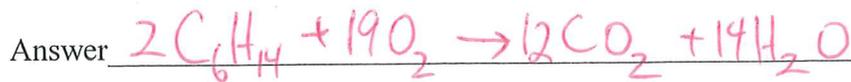


b. hydrochloric acid neutralizes aluminum hydroxide to form water & aluminum chloride



Chemistry 11

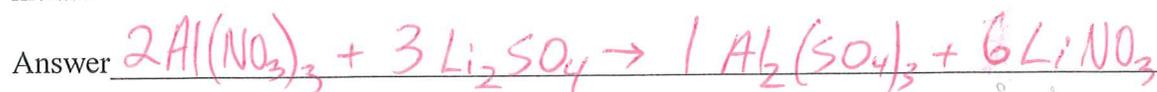
- c. hexane (C_6H_{14}) burns in oxygen to produce carbon dioxide and water



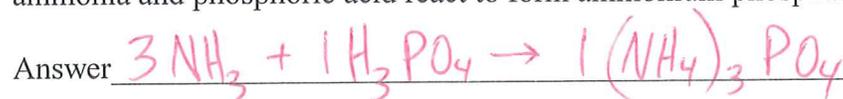
- d. carbon dioxide and water are reacted to produce glucose ($C_6H_{12}O_6$) and oxygen in photosynthesis



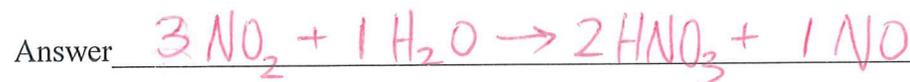
- e. aluminum nitrate reacts with lithium sulphate to form aluminum sulphate and lithium nitrate



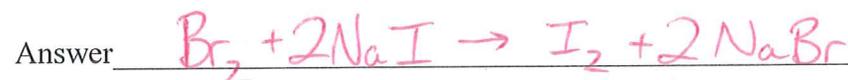
- f. ammonia and phosphoric acid react to form ammonium phosphate



- g. nitrogen dioxide reacts with water to form nitric acid and nitrogen monoxide



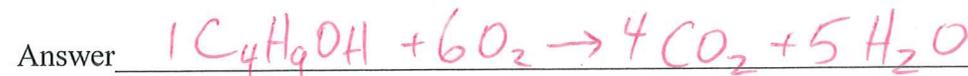
- h. bromine reacts with sodium iodide to produce iodine and sodium bromide



- i. calcium reacts in water to produce hydrogen gas and calcium hydroxide



- j. butanol (C_4H_9OH) burns in oxygen to produce carbon dioxide and water



- k. hydrogen peroxide decomposes to form water and oxygen gas



- l. aluminum bromide reacts with ammonium dichromate to produce aluminum dichromate and ammonium bromide

